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Ch 17

Thermochemistry

Practice Test Matching

Match each item with
the correct statement
below. a. calorimeter d.

enthalpy b. calorie e.

specific heat c. joule f.

heat capacity ____ 1.

quantity of heat

needed to raise the

temperature of 1 g of

water by 1°C ____ 2. SI

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unit of energy ____ 3.

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questions. This is the

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currently selected
item. Phase diagrams.
Enthalpy. Heat of
formation. Hess's law
and reaction enthalpy
change.

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Thermodynamics chemistry is the study of heat energy. The energy that is given out during chemical reactions or taken in during chemical reactions.

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The enthalpy changes, whether a reaction will take place or not. All of these different aspects are covered in thermochemistry.

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Chemistry 1: Chapter
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Test Questions. Terms
in this set (61) quantity
of heat needed to raise
the temperature of 1 g
of water by 1 degree
Celsius. calorie. SI unit
of energy. Joule.

quantity of heat
needed to change the
temperature of 1 g of a
substance by 1 degree
C. Specific heat.

Chemistry 1:
Chapter 17
Thermochemistry
Test Questions ...

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a. calorimeter d.
enthalpy b. calorie e.
Match each item with
the correct statement

below. a. heat of
reaction d. heat of
fusion b. heat of
formation e. Multiple
Choice Identify the
letter of the choice that
best completes the
statement or answers
the question.

Thermochemistry Test Preview

YES! Now is the time to

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answers. Shed the

societal and cultural

narratives holding you

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Free ...

Chapter 17

Thermochemistry183

SECTION 17.1 THE

FLOW OF

ENERGY—HEAT AND

WORK (pages 505–510)

This section explains

the relationship

between energy and

heat, and distinguishes

between heat capacity

and specific heat.

Energy

Transformations (page

505) 1. What area of

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Section
study in chemistry is
concerned with the
heat transfers that
Answers

**SECTION 17.1 THE
FLOW OF ENERGY
HEAT AND WORK
(pages 505-510)**

Thermochemistry and
Energy and
Temperature

Thermochemistry is
study of changes in
energy (heat)
associated ... notice
final answer in
problems above should

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Section

be 3 sig fig 2.09×10^4 J

or 20.9 kJ. ... Next

section will show that if

the same amount of

gas is on reactant and

product

Thermochemistry

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ISBN-13:

978-0-13252-576-3,

Publisher: Prentice Hall

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Answer the following
questions with regards
to the Heating Curve.

1) What is happening

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to the average potential energy of the molecules in the sample during section 3? Potential Energy remains constant. It only increases or decreases when temperature is remains constant on the plateaus. ...

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Review Worksheet ...

**Thermochemistry
Review Worksheet**

The Thermochemistry
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17.1 The Flow of

Energy - Sample

Problem 17.2 - Page

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17 ...

(c) This answer may be a bit confusing, until you get to the acid base section of your studies. Try not to worry about it too much.

$$\text{H}_2\text{SO}_4 + 2\text{H}_2\text{O} \rightarrow 2\text{H}_3\text{O}^+ + \text{SO}_4^{2-}$$

The protons on the sulfuric acid (an acid is a proton donor) were transferred to the water (called a proton acceptor; it is acting as a base).

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Worksheet -

Thermochemistry -

AP level

162 CHAPTER 6:

THERMOCHEMISTRY To

convert the answer to

joules, we write: 101.3

$\text{J} \cdot 0.18 \text{ L atm} \cdot 1 \text{ L atm} = -$

$\cdot \times = \cdot w -18 \text{ J} \cdot 6.17$ An

expansion implies an

increase in volume,

therefore w must be

-325 J (see the

defining equation for

pressure-volume work.)

If the system absorbs

heat, q must be $+127$

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J. The change in energy
(internal

Assessment
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CHAPTER 6

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Test :

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Name: _ Part A: Multiple
Choice. Circle the best
answer for each

question. Answer all
questions in this

section. (24 pts) 1. A
heated 250 g sample
of aluminium metal is
placed in water at

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room temperature. The water absorbs 264 J of heat from the metal and increases in temperature to 34.9°C. What

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